

LECTURE 28

- For each of the following ions, **(i)** draw a crystal field splitting diagram to show orbital occupancies in both weak and strong octahedral fields, and **(ii)** indicate the number of unpaired electrons in each case. **Label** the diagrams **(iii)** weak or strong field, **(iv)** high spin or low spin (as appropriate), **(v)** with the names of the d-orbitals, and **(vi)** with the appropriate orbital sets e_g and t_{2g} designators.
 - Fe^{2+}
 - Cr^{3+}
 - Cd^{2+}
- For each of the above ions in problem one, **(i)** Write the expected d^n electron configuration and **(ii)** calculate the CFSE (both the high and low spin states, as appropriate, and indicate pairing energies PE if electrons are paired).
- For $[\text{CoCl}_6]^{3-}$
 - Determine the coordination number of the cobalt
 - Determine the oxidation number of the cobalt
 - Predict whether this compound is high or low spin. Briefly explain your answer.
 - Estimate the octahedral crystal field splitting energy (Δ_o) in joules/mol if the wavelength most intensely absorbed is 740nm.
- Octahedral platinum(IV) complexes are used in protein crystallography to help determine three-dimensional protein structures. If the octahedral crystal field splitting energy (Δ_o) is large for these complexes,
 - Predict whether they are diamagnetic or paramagnetic
 - Write the expected d^n electron configuration
- For the following, consider a field that has a z-axis that is vertical, a y-axis that is horizontal, and x-axis that is coming out of page.
 - Draw pictures of the $d_{x^2-y^2}$ and d_z^2 orbitals
 - Predict the relative energy of $d_{x^2-y^2}$ in an octahedral crystal field compared to a linear crystal field that is along the y-axis
 - Predict the relative energy of d_z^2 in an octahedral crystal field compared to a linear crystal field that is along the y-axis
 - Predict the relative energies of $d_{x^2-y^2}$ to d_z^2 to each other in a(n)
 - octahedral field
 - linear field along the y-axis.

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